

Key

### Polyatomic Ions

- polyatomic ions most commonly end in ate, but can also end in ite or ide
- do not ever split up the atoms in polyatomic ions, they stay together as a package (think of them as a single ion)

Write the formula for the following polyatomic ions (don't forget the charge)

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|--|--|
| a) carbonate <u><math>\text{CO}_3^{2-}</math></u>        | f) hypochlorite <u><math>\text{ClO}^-</math></u>   |
| b) nitrite <u><math>\text{NO}_2^-</math></u>             | g) bisulphide <u><math>\text{HS}^-</math></u>      |
| c) oxalate <u><math>\text{C}_2\text{O}_4^{2-}</math></u> | h) hydroxide <u><math>\text{OH}^-</math></u>       |
| d) acetate <u><math>\text{CH}_3\text{COO}^-</math></u>   | I) permanganate <u><math>\text{MnO}_4^-</math></u> |
| e) cyanide <u><math>\text{CN}^-</math></u>               | j) thiocyanate <u><math>\text{SCN}^-</math></u>    |

Write the name for the following polyatomic ions (underline the ending)

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|--|--|
| a) $\text{ClO}_3^-$ <u>Chlorate</u>    | f) $\text{Cr}_2\text{O}_7^{2-}$ <u>Dichromate</u>        |
| b) $\text{HCO}_3^-$ <u>Bicarbonate</u> | g) $\text{ClO}_4^-$ <u>Perchlorate</u>                   |
| c) $\text{SO}_4^{2-}$ <u>Sulfate</u>   | h) $\text{HSO}_3^-$ <u>Bisulphite</u>                    |
| d) $\text{NO}_3^-$ <u>Nitrate</u>      | I) $\text{HC}_2\text{O}_4^-$ <u>Binoxalate</u>           |
| e) $\text{PO}_4^{3-}$ <u>Phosphate</u> | j) $\text{H}_2\text{PO}_4^-$ <u>Dihydrogen phosphate</u> |