

**Atomic Structure (Sub-Atomic
Particles) Worksheet**

Name: Key

Name	Isotope Notation	Atomic Number	Atomic Mass	Protons	Neutrons	Electrons	Net Charge
Fluoride atom	${}^{19}_9\text{F}$	9	19	9	10	9	0
Zn atom	${}^{65}_{30}\text{Zn}$	30	65	30	35	30	0
Potassium atom	${}^{39}_{19}\text{K}$	19	39	19	20	19	0
K ion	${}^{40}_{19}\text{K}^{1+}$	19	40	19	21	18	+1
P atom	${}^{31}_{15}\text{P}$	15	31	15	16	15	0
Cu atom	${}^{63}_{29}\text{Cu}$	29	63	29	34	29	0
Fe atom	${}^{56}_{26}\text{Fe}$	26	56	26	30	26	0
U ion	${}^{235}_{92}\text{U}^{2+}$	92	235	92	143	90	+2
U atom	${}^{238}_{92}\text{U}$	92	238	92	146	92	0
Mg ion	${}^{24}_{12}\text{Mg}^{2+}$	12	24	12	14	10	+2
Chloride ion	${}^{37}_{17}\text{Cl}^{-}$	17	37	17	20	18	-1
Chloride ion	${}^{35}_{17}\text{Cl}^{-}$	17	35	17	18	18	-1
Sr ion	${}^{88}_{38}\text{Sr}^{+2}$	38	88	38	50	36	+2
Al ion	${}^{27}_{13}\text{Al}^{3+}$	13	27	13	14	10	+3
Copper (II) ion	Cu^{2+}	29	65	29	36	27	+2
Boron atom	${}^{11}_5\text{B}$	5	11	5	6	5	0
Nitrogen ion	${}^{14}_7\text{N}^{3-}$	7	14	7	7	10	-3
Pb ion	${}^{207}_{82}\text{Pb}^{4+}$	82	207	82	125	78	+4
Fe ion	${}^{56}_{26}\text{Fe}^{2+}$	26	56	26	30	24	+2
Sn ion	${}^{119}_{50}\text{Sn}^{2+}$	50	119	50	69	48	+2
Bromide ion	${}^{80}_{35}\text{Br}^{-}$	35	80	35	45	36	-1
As ion	${}^{75}_{33}\text{As}^{3-}$	33	75	33	42	36	-3
W atom	${}^{184}_{74}\text{W}$	74	184	74	110	74	0
He atom	${}^4_2\text{He}$	2	4	2	2	2	0
Na ion	${}^{23}_{11}\text{Na}^{+}$	11	23	11	12	10	+1
Al ion	${}^{26}_{13}\text{Al}^{3+}$	13	26	13	13	10	+3
Cs atom	${}^{133}_{55}\text{Cs}$	55	133	55	78	55	0
Cr ion	${}^{52}_{24}\text{Cr}^{3+}$	24	52	24	29	21	+3