Key (only Net Ionic)

NET IONIC EQUATIONS

For each of the following write a ***molecular (formula) equation***, a ***complete (total) ionic equation***, and a ***net ionic equation*** (DO NOT FORGET YOUR SUBSCRIPTS INDICATING PHASE): ** I would suggest using a separate sheet for your answers.

1. Barium chloride and Strontium nitrate — No net ionic equation
2. Barium chloride and Aluminum sulfate — $Ba^{2+}_{(aq)} + SO_4^{2-}_{(aq)} \longrightarrow BaSO_{4(s)}$
3. Barium chloride and Silver nitrate — $Ag^{+}_{(aq)} + Cl^{-}_{(aq)} \longrightarrow AgCl_{(s)}$
4. Strontium nitrate and Aluminum sulfate — $Sr^{2+}_{(aq)} + SO_4^{2-}_{(aq)} \longrightarrow SrSO_{4(s)}$
5. Strontium hydroxide and Silver nitrate — $Ag^{+}_{(aq)} + OH^{-}_{(aq)} \longrightarrow AgOH_{(s)}$
6. Aluminum sulfate and Silver nitrate — $2Ag^{+}_{(aq)} + SO_4^{2-}_{(aq)} \longrightarrow Ag_2SO_{4(s)}$
7. Nickel (II) chloride and Magnesium bromide — No net ionic equation
8. Nickel (II) chloride and Sodium sulfate — No net ionic equation
9. Nickel (II) chloride and Sodium phosphate — $3Ni^{2+}_{(aq)} + 2PO_4^{3-}_{(aq)} \longrightarrow Ni_3(PO_4)_{2(s)}$
10. Copper (II) sulfate and Potassium chloride — No net ionic equation
11. ~~Zinc~~ Calcium iodide and Copper (I) sulfate — $Ca^{2+}_{(aq)} + 2I^{-}_{(aq)} + 2Cu^{+}_{(aq)} + SO_4^{2-} \longrightarrow CaSO_{4(s)} + 2CuI_{(s)}$
12. Barium bromide and Sodium sulfate — $Ba^{2+}_{(aq)} + SO_4^{2-}_{(aq)} \longrightarrow BaSO_{4(s)}$
13. Magnesium chloride and Sodium hydroxide — $Mg^{2+}_{(aq)} + 2OH^{-}_{(aq)} \longrightarrow Mg(OH)_{2(s)}$
14. Barium nitrate and Potassium sulfide — No net ionic equation.
15. Calcium chloride and Potassium hydroxide — $Ca^{2+}_{(aq)} + 2OH^{-}_{(aq)} \longrightarrow Ca(OH)_{2(s)}$
16. Sodium sulfate and Lead (II) nitrate — $Pb^{2+}_{(aq)} + SO_4^{2-}_{(aq)} \longrightarrow PbSO_{4(s)}$
17. Sodium sulfate and Barium iodide — $Ba^{2+}_{(aq)} + SO_4^{2-}_{(aq)} \longrightarrow BaSO_{4(s)}$
18. Potassium sulfate and Strontium nitrate — $Sr^{2+}_{(aq)} + SO_4^{2-}_{(aq)} \longrightarrow SrSO_{4(s)}$
19. Sodium hydroxide and Ammonium sulfide — No net ionic equation.
20. Strontium hydroxide and Copper (II) sulfate

$$Sr^{2+}_{(aq)} + 2OH^{-}_{(aq)} + Cu^{2+}_{(aq)} + SO_4^{2-}_{(aq)} \longrightarrow SrSO_{4(s)} + Cu(OH)_{2(s)}$$